

The H_3O^+ concentration in an aqueous solution is a quantitative measure of acidity the higher the H_3O^+ concentration, the more acidic the solution. In contrast the higher the concentration of OH^- the more basic the solution.

3. Discuss in detail how different solutions in an aqueous system exhibit increase in acidity and an increase in basicity based on H^+ and OH^- ?

Ans: The concept of pH was first introduced by the Danish chemist Soren Peder Lauritz Sorensen at the Carlsberg Laboratory in 1909. "pH" is an abbreviation for "potential hydrogen" and is a scale used for measuring the relative acidity or alkalinity of a Liquid solution.

pH: A measurement of the concentration of Hydrogen ions (H^+) in a solution. It may also be defined as "the negative Logarithm of Hydrogen ion concentration".

Mathematically $\text{pH} = -\log [\text{H}^+]$

pOH: A measurement of the concentration of Hydroxyl ions (OH^-) in a solution. It may also be defined as "the negative logarithm of Hydroxyl ion concentration".

Mathematically $\text{pOH} = -\log [\text{OH}^-]$

pH Scale: The pH scale measures how acidic or basic a substance. The scale has values ranging from zero (the most acidic) to 14 (the most basic). The pH of different solutions can be measured by pH paper, pH indicator and pH meter.

4. What do you mean by balancing of neutralization reaction with the help of examples?

Ans: **Balancing of Neutralization Reaction:**

The balancing of a neutralization reaction involves the equal consumption of H^+ and OH^- in an aqueous solution due to this we have to look at the acidity of the base and basicity of acid in the balancing equation

For example, the chemical equation of neutralization of sodium hydroxide and hydrochloric acid is as follows:



NaOH and HCl are neutralizing each other in a ratio of 1:1 because one mole of NaOH releases one mole of OH^- and one mole of HCl releases one mole of H^+ . Similarly, NaOH or KOH when reacted with HBr, HI and HNO_3 shows a ratio of 1:1. One mole of each HI, HNO_3 and Hbr releases one mole of H^+ :

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Marye Anne Fox, James K. Whitesell



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