

# ATOMIC STRUCTURE

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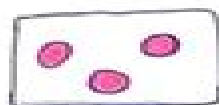
**ATOM** - The basic building block for all matter in the universe. Single particle.

**PROTON** - A positively charged particle that is found in the nucleus of the atom.

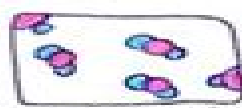
**NEUTRON** - A particle found in the nucleus with no charge.

**ELECTRON** - A negatively charged particle that spins around the outside of the nucleus.

**ELEMENT** - A substance consisting of atoms which all have the same number of protons.



ELEMENT



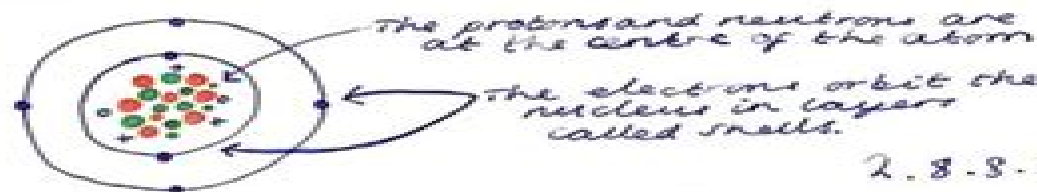
COMPOUND



MIXTURE

**COMPOUND** - 2 or more different elements put together, chemically joined together.

**MIXTURE** - combination of two or more pure substances.



**ATOMIC NUMBER** = number of protons.

Atomic mass - neutron = proton / electron.

2.8.8.13

## THE LAW OF CONSERVATION OF MASS

The total mass of the products formed in a reaction is equal to the total mass of the reaction.

[Mass number = number of protons + number of neutrons.]

**IONS** - an ion is an atom or group of atoms that have an electrical charge, either positive or negative.

• Elements can gain or lose electrons during a chemical reaction.

• **POSITIVE ION** - an atom that loses electrons has more protons than electrons and so has a positive overall charge.

• **NEGATIVE ION** - an atom that gains electrons has more electrons than protons and so has a negative overall charge.

- protons don't change.

• **METALS** tend to [lose] electrons.

• **NON-METALS** tend to [gain] electrons.

• **ISOTOPES** - are atoms that have the same number of protons but different numbers of neutrons in their nuclei.

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1. Understanding the eBook Gcse Chemistry Revision Guide S
  - The Rise of Digital Reading Gcse Chemistry Revision Guide S
  - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Chemistry Revision Guide S
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gcse Chemistry Revision Guide S
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Chemistry Revision Guide S
  - Personalized Recommendations
  - Gcse Chemistry Revision Guide S User Reviews and Ratings
  - Gcse Chemistry Revision Guide S and Bestseller Lists
5. Accessing Gcse Chemistry Revision Guide S Free and Paid eBooks
  - Gcse Chemistry Revision Guide S Public Domain eBooks
  - Gcse Chemistry Revision Guide S eBook Subscription Services
  - Gcse Chemistry Revision Guide S Budget-Friendly Options
6. Navigating Gcse Chemistry Revision Guide S eBook Formats

- ePub, PDF, MOBI, and More
  - Gcse Chemistry Revision Guide S Compatibility with Devices
  - Gcse Chemistry Revision Guide S Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Gcse Chemistry Revision Guide S
    - Highlighting and Note-Taking Gcse Chemistry Revision Guide S
    - Interactive Elements Gcse Chemistry Revision Guide S
  8. Staying Engaged with Gcse Chemistry Revision Guide S
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Gcse Chemistry Revision Guide S
  9. Balancing eBooks and Physical Books Gcse Chemistry Revision Guide S
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Gcse Chemistry Revision Guide S
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Gcse Chemistry Revision Guide S
    - Setting Reading Goals Gcse Chemistry Revision Guide S
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Gcse Chemistry Revision Guide S
    - Fact-Checking eBook Content of Gcse Chemistry Revision Guide S
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
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