

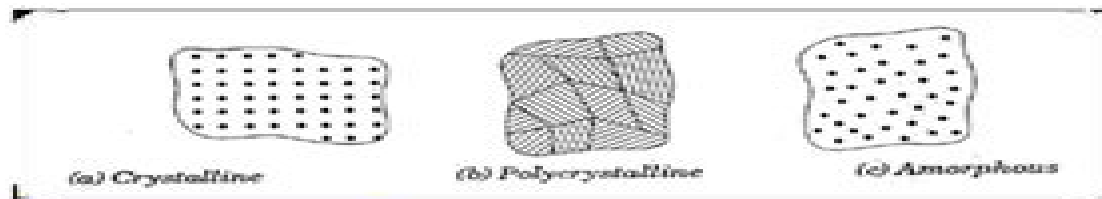
CRYSTALLOGRAPHY AND CRYSTAL STRUCTURES

CRYSTALLOGRAPHY

The branch of science which deals with the study of geometric form and other physical properties of the crystalline solids by using X-rays, electron beam, and neutron beams etc is called crystallography or crystal physics.

The solids are classified into two types **crystalline and amorphous**.

- A substance is said to be crystalline, when the arrangement of atoms, molecules or ions inside it is regular and periodic. Ex. NaCl, Quartz crystal. Though two crystals of same substance may look different in external appearance, the angles between the corresponding faces are always the same.
- In amorphous solids, there is no particular order in the arrangement of their constituent particles. Ex. Glass.



CRYSTALLINE SOLIDS

1. Crystalline solids have regular periodic Arrangement of particles (atoms, ions, Or molecules).
2. They are un-isotropic i.e.; they differ in properties with direction.
3. They have well defined melting and freezing points. Melting and freezing points occurs at different temperatures at different locations in the solids.
4. Crystalline solids may be made up of materials are metallic crystals or non-metallic crystals. Some of the metallic crystals are Copper, silver, aluminium, tungsten, and manganese. Non-metallic crystals are crystalline carbon, crystallized polymers or plastics.
5. Metallic crystals have wide use in engineering because of their favourable Properties of strength, ductility, conductivity and reflection.

Lecture Notes On Crystallography 3ed

J Dewey



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