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**Jim Elferdink, David Reynolds**



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web vocabulary solute solvent solution solubility saturated unsaturated supersaturated independent variable dependent variable use a vocabulary term to answer the first two questions below 1 which term applies to a sponge that is dry 2 which term applies to a sponge that is soaked 3

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web chem 116 pogil worksheet week 4 properties of solutions key questions 1 identify the principal type of solute solvent interaction that is responsible for forming the following solutions a kno 32 in water b br in benzene c 6 h 6 c glycerol ch 22 oh ch oh ch oh in water d hcl in acetonitrile ch 3 cn hcl does not form ions in

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web model 1 three solutions the following data refer to three experiments in which solute is added to water in a beaker at 20 oc the mixtures are stirred and then allowed to sit for three hours before measuring the amount of solid that dissolves ten

separate trials are conducted for each experiment

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web pogil principles of solubility why the previous pogil discussed a few factors affecting how and why solutions form based on entropy and interparticle forces another factor affecting solution formation is the enthalpy of the solution or the energy required for the solution to form

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web jul 12 2023 the common ion effect of  $\text{H}_3\text{O}^+$  on the ionization of acetic acid the common ion effect suppresses the ionization of a weak base by adding more of an ion that is a product of this equilibrium example 18 3 4 consider the common ion effect of  $\text{OH}^-$  on the ionization of ammonia

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web if a solute is soluble in a solvent the overall process is said to be spontaneous if a solute does not dissolve in a solvent insoluble the solution process is said to be non spontaneous a spontaneous process or change proceeds in a given direction without needing to be driven by an external energy input

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web solubility and solubility product mh 5 16 1 16 2 in this section we are going to consider the solubility of ionic solids in water solubility may be considered to be an equilibrium the equilibrium is between solid and ions in solution any ionic solid is 100% ionized in aqueous solution once it actually dissolves

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web information solubility rules a solute is considered soluble if an appreciable amount of it can be dissolved in a given amount of the solvent for example both table salt  $\text{NaCl}$  and table sugar  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$  are soluble substances in water a solute is considered insoluble if very little of it dissolves in a given amount of the solvent

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web predict what might happen to the solubility of the insoluble salt if a 1.0 M acid solution was added drop wise to the beaker hint which ion species in the beaker will react with the acid how will this reaction alter the concentration of the common ion