

# Sentence Structures

## Simple

A simple sentence is a sentence that has one subject and one verb, and it expresses a complete idea. It can stand alone as a complete thought.

**Examples:** The cat purrs. Birds fly.

## Compound

A compound sentence is a sentence that joins two simple sentences together. It uses words like "and," "but," or "or" to connect these ideas.

**Examples:** I like ice cream, but I don't like cake.

## Complex

A complex sentence is a sentence that joins a simple sentence (a complete thought) with an extra part that can't stand alone.

**Examples:** I like to read because it helps me learn new things.

## Compound-Complex

A compound-complex sentence is like a mix of different kinds of sentences. It has more than one action and also includes extra information.

**Examples:** "I like to read, and when it rains, I stay indoors."



# English Sentence Structure

**Phil Williams**



## **English Sentence Structure :**

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English Sentence Structure Robert Krohn, University of Michigan. English Language Institute, 1971 For ESL students at an intermediate level

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**English Sentence Structure** Thomas Cyprian Jupp, 1979

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup> eg 2 HNO<sub>3</sub>(aq) + KOH(aq) → KNO<sub>3</sub>(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup>

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produced when one mole of hcl or one mole of naoh reacts